

## Chemistry Unit 4 Outline: The Mole and Chemical Composition

### Chapter 7: The Mole and Chemical Composition

Classes	Topics	Suggested Reading	✓	Assignments	✓
1	Mole, Avogadro's Number, Molar Mass, Converting between Mass, Moles, and Number of Particles using Molar Mass and Avogadro's Number, Calculating Average Atomic Mass	7.1 Avogadro's Number and Molar Conversions (pg. 224 – 233)  7.2 Relative Atomic Mass and Chemical Formulas (pg. 234 – 233)		pg. 228 #1 to 4 (Practice) pg. 229 #1 to 5 (Practice) pg. 231 #1 to 4 (Practice) pg. 232 #1 to 3 (Practice) pg. 233 #1 to 13  pg. 239–240 #1 to 4 (Practice) pg. 236 #1 and 2 (Practice) pg. 240 #1, 3 to 12, 14 to 16	
2	Percentage Composition, Empirical Formula, Molecular Formula, Calculating Percentage Composition from Chemical Formulas, Determining Empirical and Molecular Formulas from Percentage Composition and Molar Mass	7.3 Formulas and Percentage Composition (pg. 241 – 248)		pg. 243 #1 to 4 (Practice) pg. 245 #1 to 3 (Practice) pg. 248 #1 to 5 (Practice) pg. 248 #1 to 10	
3	<b>Lab #4: Percentage Composition of Hydrates</b> <b>(A Block: December 13, Tues)</b> <b>(C, D &amp; E Blocks: December 12, Mon)</b>	Lab #4 Procedure		<b>Lab #4 Report Due</b> <b>(A Block: Dec 15, Thurs)</b> <b>(C, D, &amp; E Block: Dec 14, Wed)</b>	
4	Chapter 7 Review			pg. 251–254 #19 to 66	
5	<b>Unit 4 Test with Final Exam</b> <b>(A Block: December 16, Fri)</b> <b>(C &amp; D Block: December 19, Mon)</b> <b>(E Block: December 20, Tues)</b>				