

## Honour Chemistry Unit 1 Outline: Basic Chemistry

### Chapter 1: Introduction

Classes	Topics	Suggested Reading	✓	Assignments	✓
1	Course Outline, Lab Writeup				
2	Lab Safety, Scientific Method, Theory, Scientific Law, Accuracy, Reliability, Classification of Matter (Pure Substance and Mixture, Elements and Compounds), Three States of Matter, Physical, Chemical, Extensive and Intensive Properties, Mass, Volume, Density	Lab Safety Contract and Video 1.1 The Study of Chemistry (pg. 2) 1.2 The Scientific Method (pg. 2 – 3) 1.3 Classification of Matter (pg. 4 – 7) 1.4 Physical and Chemical Properties of Matter (pg. 7 – 8)		pg. 23 #2 to 12 (even)	
2	Lab Safety Video, Lab Walk-Through, Common Laboratory Equipment & Apparatus, Common Lab Techniques (Meniscus, Lighting Bunsen Burner, Electronic Balance, Pipetting)	Lab Safety, Equipment and Techniques Handout		Read & Sign Lab Safety Contract	
3	Lab Safety Video and Lab Walk-Through Microscopic and Macroscopic Properties, SI Units, Scientific (Exponential) Notations, Precision, Percent Error and Percent Yield, Exact Numbers, Uncertainty, Significant Digits, Mass vs. Weight, Temperature Scales, Density Calculations, Unit Factor (Analysis) Method	1.5 Measurement (pg. 8 – 12) 1.6 Handling Numbers (pg. 13 – 18) 1.7 Dimensional Analysis in Solving Problems (pg. 18 – 22)		pg. 23–24 #14, 16, 18 to 20, 22, 24 to 26, 28 to 30  pg. 24 #32 to 34, 36, 38, 40	
4	<b>Chapter 1 Quiz and Lab Safety Quiz (September 9, Wed)</b>				

### Chapter 2: Atoms, Molecules and Ions

Classes	Topics	Suggested Reading	✓	Assignments	✓
1 & 2	History of Chemistry, Law of Conservation of Mass, Law of Definite and Multiple Proportions, Atomic Theories (Dalton, J.J Thomson, Nuclear, Quantum), Millikan Oil drop Experiment, Atomic Structures and Subatomic Particles, Radioactivity and Particles, Isotopes, Atomic Number and Mass Number, Valence Electrons, Atomic Orbitals, Periodic Table of Elements (Mendeleev), Molecules, Ions (Cations and Anions), Metals, Non-Metals, Ionic and Covalent Compounds, Groups (Families), Periods, Binary and Polyatomic Elements and Molecules	2.1 The Atomic Theory (pg. 29 – 30) 2.2 The Structure of Atom (pg. 30 – 35) 2.3 Atomic Number, Mass Number and Isotopes (pg. 35 – 36) 2.5 Molecules and Ions (pg. 38 – 39)  2.4 The Periodic Table (pg. 36 – 38)		pg. 53 #4, 8 pg. 54 #10 to 16 (even)  pg. 54–55 #26 to 34 (even) pg. 54 #18 to 24 (even)	
3 & 4	Chemical and Molecular Formulas, Allotrope, Structural Formula, Empirical Formulas, Nomenclature of Ionic Compounds, Polyatomic Ions and Oxoanions, Nomenclature of Hydrates, Nomenclature of Molecular Compounds, Names and Formulas of some Common Molecular Compounds, Nomenclature of Acids	2.6 Chemical Formulas (pg. 39 – 43) 2.7 Naming Compounds (pg. 43 – 51) 2.8 Introduction to Organic Compounds (pg. 51 – 52) <i>(Read only: won't be tested until Unit 5)</i>		pg. 55 #36, 38, 42, 44, 46 pg. 55 #47 to 50	
5	<b>Lab #1: Lab Safety, Measuring Techniques &amp; Diagnostic Tests for H<sub>2</sub>, O<sub>2</sub>, and CO<sub>2</sub> (September 16, Wednesday)</b>	Lab #1 Procedure		<b>Lab #1 Report Due (September 24, Thurs)</b>	
6	<b>Chapter 2 Quiz (September 17, Thurs)</b>				

### Chapter 3: Stoichiometry

Classes	Topics	Suggested Reading	✓	Assignments	✓
1	Atomic Mass., Atomic Mass Unit (amu), Average Atomic Mass and Relative Abundance, Mole, Avogadro's Number, Molar Mass, Conversions between Mass, Mole, and Molar Mass $\left(n = \frac{m}{M}\right)$ , Mass Spectrometer	3.1 Atomic Masses (pg. 59 – 60) 3.2 Avogadro's Number and Molar Mass of an Element (pg. 60 – 64) 3.3 Molecular Mass (pg. 64 – 66) 3.4 The Mass Spectrometer (pg. 66 – 67)		pg. 86 #4, 6, 8 pg. 86–87 #10, 13 to 16, 18 to 20, 22  pg. 87 #24 to 30 (even)	
2	Mass Percent, Determination of Empirical Formulas and Molecular Formulas and Hydrates, Chemical Reactions, Reactants, Products, Coefficients, Classifying and Balancing Different Types of Chemical Reactions (Composition, Decomposition, Single and Double Replacements, Hydrocarbon Combustion)	3.5 Percent Composition of Compounds (pg. 67 – 70) 3.6 Experimental Determination of Empirical Formulas (pg. 70 – 73) 3.7 Chemical Reactions and Chemical Equations (pg. 73 – 77)		pg. 87–88 #40 to 54 (even)  pg. 88 #59, 60	
3	<b>Chapter 3 Quiz (covers 3.1 to 3.7) (September 30, Wednesday)</b>				
4 & 5	Stoichiometry, Mole Ratio, Gravimetric Stoichiometry, Excess and Limiting Reagents Calculations, Reaction Yield	3.8 Amounts of Reactants and Products (pg. 77 – 81) 3.9 Limiting Reagent (pg. 81 – 83) 3.10 Reaction Yield (pg. 83 – 85)		pg. 89–90 #64 to 78 (even) pg. 90 #82, 84, 86 pg. 91 #90, 92	
6	<b>Unit 1 Review</b>			Unit 1 Practice Test	
7	<b>Unit 1 Test (October 6, Tuesday)</b>				